

Analytical Mechanics of Chemical Reactions. V. Application to the Linear Reactive $H+H_2$ Systems*

SHIOU-FU WU AND R. A. MARCUS

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Action-angle variables, which were used to treat atomic and molecular structure many years ago,⁶ have been used more recently in this laboratory to treat inelastic and reactive collisions.^{2,7-9} Their use led to relatively simple approximate expressions for energy transfer and other properties. Encouraging agreement with exact results was obtained for the inelastic systems.^{7,8} The method provided information about alternative quantum approximations (near static vs near adiabatic), because of the close relationship which the Hamilton-Jacobi equation bears to the Schrödinger equation and which the action variables J_i bear to the quantum numbers n_i ($J_i/h \cong n_i + \frac{1}{2}$ or n_i , depending upon coordinate). The approximations are particularly suited to near-elastic or near-adiabatic systems.

The rearrangement reaction $H+H_2 \rightarrow H_2+H$ has been intensively studied both classically¹⁰⁻¹³ and quantum mechanically,^{14,15,16-18} largely though not exclusively numerically. The present paper calculates and compares, for the first time, nonadiabatic corrections with exact numerical results. Some quantum mechanical implications are noted in a concluding section. Results on reaction in a plane will be given in a subsequent paper.

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$$\dot{p} = \frac{\rho(d\kappa/ds)p_s^2}{\mu\eta^3} - \frac{\partial V}{\partial s}, \quad (1a)$$

$$\dot{s} = p_s/\mu\eta^2, \quad (1b)$$

$$\dot{p}_\rho = (\kappa p_s^2/\mu\eta^3) - (\partial V/\partial \rho), \quad (1c)$$

$$\dot{\rho} = p_\rho/\mu, \quad (1d)$$

where the p 's denote momenta conjugate to s and ρ and where

$$\eta = 1 + \kappa\rho. \quad (2)$$

The force components in Eqs. (1) along the natural collision coordinates, $-\partial V/\partial \rho$ and $-\partial V/\partial s$, are obtained from the interatomic forces using the "chain rule,"

$$\frac{\partial V}{\partial q^i} = \sum_{k=1}^3 \sum_{j=1}^2 \frac{\partial V}{\partial R_k} \frac{\partial R_k}{\partial q^i} \frac{\partial Q^j}{\partial q^i}, \quad (3)$$

where $q^i = s, \rho$, $Q^j = z, Z$, and R_1, R_2 , and R_3 are the internuclear distances. The four simultaneous differential equations (1) can be integrated numerically.

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The potential energy function $V(\rho, s)$ is written² in terms of its value on curve C , $V_1(s)$, and the increment to a finite ρ at this s , $V_2(\rho, s)$,

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the α/η^2 term in the zeroth-order treatment of the ρ motion, and we do this in the present case²:

The effective potential $U(\rho, s)$ for the ρ motion when α/η^2 is included is²

$$U(\rho, s) = V_2(\rho, s) + (\alpha/\eta^2), \quad (5)$$

where $\alpha(s)$ denotes the average $p_s^2/2\mu$ in the adiabatic approximation¹⁸:

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where

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The $\rho_0(s)$ and the adiabatic vibrational energy $[E_{\text{vib}}^{\text{ad}}(s)]$ in (6) are given below. $\alpha(s)$ is found by solving (6), (7), and (9) iteratively or simultaneously. [An iteration can be initiated by setting $V_2(\rho_0, s) = 0$ in (6) and $\partial^2 U/\partial \rho^2 = \partial^2 V/\partial \rho^2$ in (9) in the first step.]

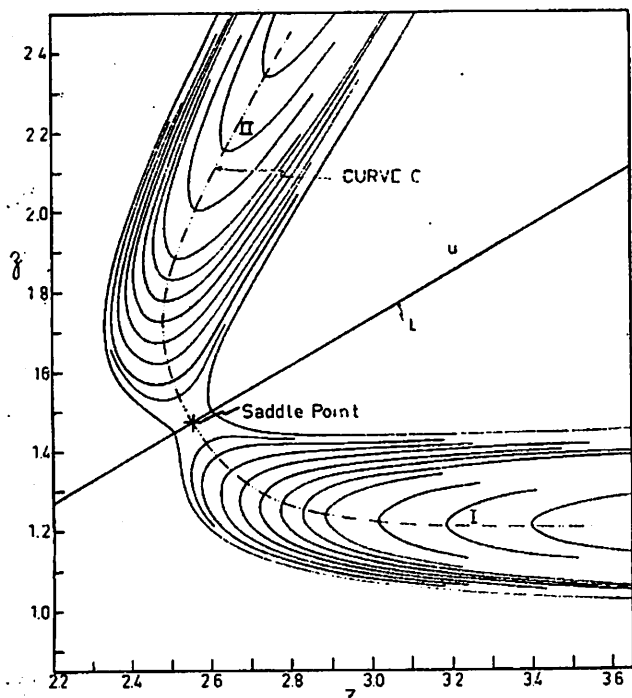


FIG. 1. Potential energy contour map for H_3 (relative to H , H_2 as zero) in the (z, Z) plane for linear configurations. Curve C (---) is chosen graphically as the curve which is steepest descent from the saddle point down to the reactant and product valleys, respectively.

The value of ρ which minimizes $U(\rho, s)$ at any s is denoted² by $\rho_0(s)$,

$$\partial U(\rho, s)/\partial \rho = 0 \quad \text{at} \quad \rho = \rho_0(s). \quad (7)$$

The vibrational variables ρ and p_ρ can be expressed, at any s , in terms of local action-angle variables for the ρ motion, J_ρ and w_ρ , respectively.² In physical systems the initial value of w_ρ is random in the interval $(0, 1)$. In the adiabatic approximation J_ρ is a constant of the motion. When the ρ motion is treated as locally harmonic, we have²

$$U(\rho, s) \cong U(\rho_0, s) + \frac{1}{2}\mu\omega^2(\rho - \rho_0)^2, \quad (8)$$

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1.2342	2.9351	0.008975
1.2261	3.0186	0.008091
1.2214	3.1096	0.007239
1.2191	3.1838	0.006620
1.2140	3.3988	0.005147
1.2138	3.5911	0.004149
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1.2134	4.2314	0.002150

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$$\dot{J}_\rho = -2\dot{\rho}_0(\mu\omega\pi J_\rho)^{1/2} \sin 2\pi w_\rho - (J_\rho \dot{\omega}/\omega) \cos 4\pi w_\rho, \quad (11)$$

$$\dot{w}_\rho = (\omega/2\pi) - \frac{1}{2}\dot{\rho}_0(\mu\omega/J_\rho\pi)^{1/2} \cos 2\pi w_\rho + (\dot{\omega}/4\pi\omega) \sin 4\pi w_\rho. \quad (12)$$

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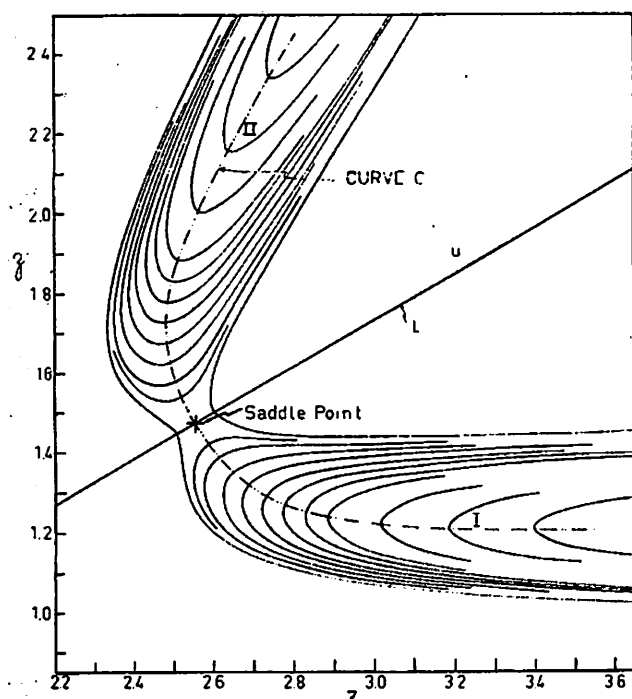


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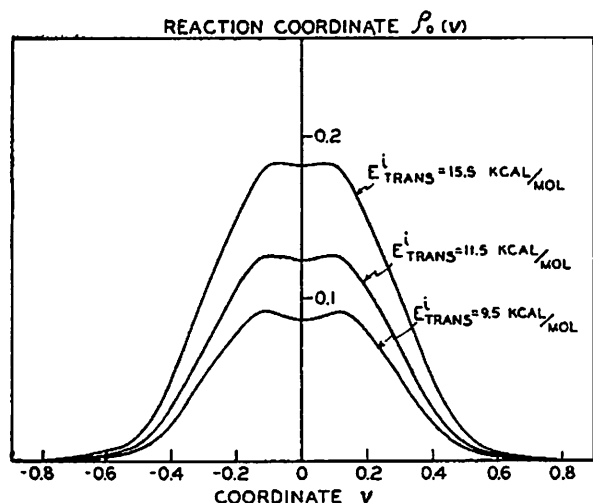


FIG. 2. Reaction coordinate $\rho_0(v)$ as a function of coordinate v for three initial relative translational energies, $E_{\text{TRANS}}^i = 9.5$, 11.5, and 15.5 kcal/mol.

variables from (J_ρ, w_ρ) to canonically conjugate variables a and a^* can be performed,² where a is $J_\rho^{1/2} \exp(-2\pi i w_\rho)$, yielding²

$$\dot{a} = -i\omega a + i\dot{\rho}_0(\mu\omega\pi)^{1/2} - (\dot{\omega}a^*/2\omega). \quad (13)$$

a^* is the complex conjugate of a .

When the $\dot{\omega}$ term in (13) is treated as a perturbation, integration of (13) yields (cf. Eqs. 93 and 94 of Paper IV)¹⁹

$$J_\rho = J_\rho^0 - 2(J_\rho^0)^{1/2} \int_{-\infty}^t \left[\dot{\rho}_0(\mu\omega\pi)^{1/2} \times \sin \left(\int_0^{t'} \omega dt'' + \delta_\rho \right) \right] dt' + \left| \int_{-\infty}^t \left[\dot{\rho}_0(\mu\omega\pi)^{1/2} \times \exp \left(i \int_0^{t'} \omega dt'' \right) \right] dt' \right|^2 + \Delta, \quad (14)$$

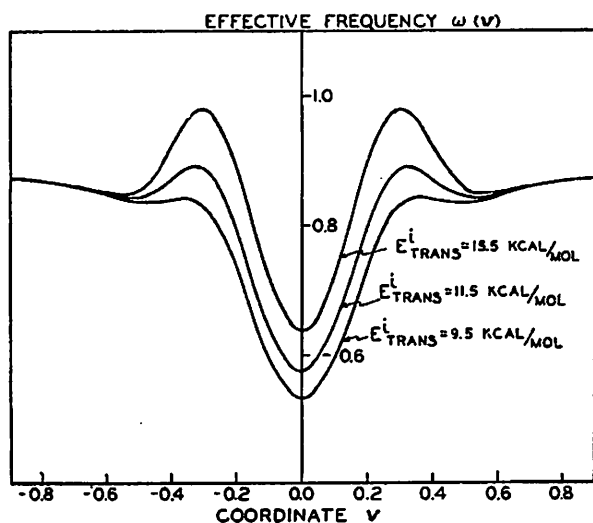


FIG. 3. Effective frequency $\omega(v)$ of the vibrational ρ motion as a function of coordinate v for three initial relative translational energies, $E_{\text{TRANS}}^i = 9.5$, 11.5, and 15.5 kcal/mol.

where Δ is the perturbation due to the $\dot{\omega}$ term:

$$\Delta = -J_\rho^0 \int_{-\infty}^t \left[\frac{\dot{\omega}}{\omega} \cos 2 \left(\int_0^{t'} \omega dt'' + \delta_\rho \right) \right] dt'. \quad (15)$$

When both the $\dot{\rho}_0$ and $\dot{\omega}$ terms in (13) are treated as perturbations to first order, Eq. (14) is again obtained,² apart from the δ_ρ -independent term (the third term) on the right-hand side of (14). In effect, therefore, Eq. (14) treats the $\dot{\omega}$ term to first order and the $\dot{\rho}_0$ term to second or higher order.

APPLICATION TO THE $\text{H} + \text{H}_2 \rightarrow \text{H}_2 + \text{H}$ REACTION

A. Potential Energy Surface, Curve C, and Other Properties

A number of studies have been made on obtaining the best potential energy surface for H_3 system.²⁰ We employ the Porter and Karplus surface,²¹ previously used in extensive trajectory¹¹ and other scattering calculations.¹⁶ The contour of the potential energy surface expressed in the skewed-coordinates is plotted in Fig. 1.

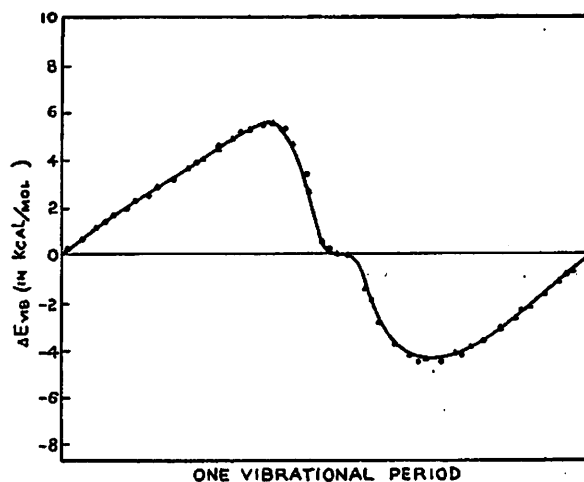


FIG. 4. Comparison of exact calculations for ΔE_{vib} 's based on natural collision coordinates (●) and on ordinary coordinates (▲); linear collision of $\text{H} + \text{H}_2$ system having $E_{\text{TRANS}}^i = 9.5$ kcal/mol.

A curve C can be obtained from the contours, for example, by drawing the curve of the steepest descent from the saddle point. For numerical purposes this curve (the points of which are given in Table I) was fitted to a polynomial by least-square method. If the (z, Z) axes of Fig. 1 are rotated by 30° about the saddle-point into (u, v) axes to utilize the symmetry of the surface, the polynomial equation describing curve C is, in these (u, v) coordinates,

$$\begin{aligned} u &= -(3v)^{1/2} + 2r_0 & \text{for } v < -0.9 \text{ a.u.,} \\ u &= (3v)^{1/2} + 2r_0 & \text{for } v > 0.9 \text{ a.u.,} \\ u &= \sum_{l=0}^6 A_l v^{2l} & \text{for } -0.9 \leq v \leq 0.9 \text{ a.u.,} \end{aligned} \quad (16)$$

where r_0 is the equilibrium bond length (in reduced, mass-weighted units) and where the coefficients A_i are given in Table II. The saddle-point occurs at $u=0, v=0$.

Using this curve C , $\rho_0(s)$ and $\omega(s)$ were calculated from Eqs. (7) and (9) for various energies, and are shown for several energies in Figs. 2 and 3. As seen from these plots both quantities have already effectively approached their asymptotic values at $|v|=0.9$ a.u. Thus, it is a good approximation to truncate the calculation at this $|v|$.

Both $\rho_0(s)$ and $\omega(s)$ were fitted to a polynomial in v by the least-square method to obtain an analytical expression for these quantities.

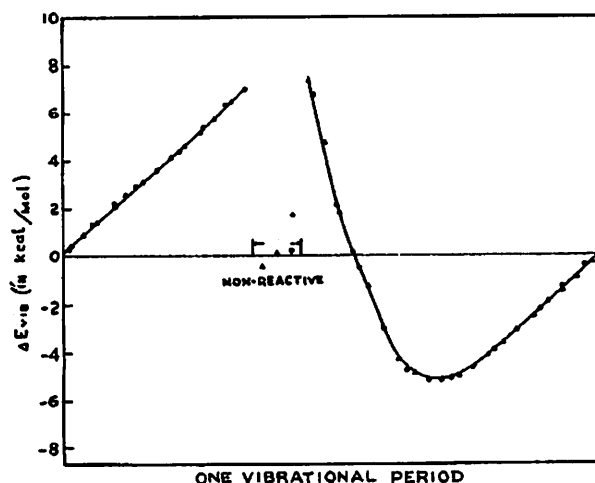


FIG. 5. Comparison of exact calculations for ΔE_{vib} 's based on natural collision coordinates ($\bullet$) and on ordinary coordinates ($\blacktriangle$); linear collision of $H+H_2$ system having $E_{trans}^i=10.55$ kcal/mol).

B. Comparison of Exact Calculations Based on Eqs. (1) with Those Based on Cartesian Coordinates, z and Z

Before comparing the results based on Eqs. (14) with those based on exact (numerical) integration in ordinary center-of-mass coordinates, it is useful to compare the results of the latter with those based on the exact integration of (1). The results of these purely numerical integrations are compared as a function of initial vibrational phase in Figs. 4 and 5. To obtain these figures the equations of motion in each set of coordinates [Eq. (1) for natural collision coordinates and the usual equations for ordinary coordinates²²] were integrated by a fourth-order Runge-Kutta-Gill method with the appropriate boundary conditions.²³ A total number of 36 vibrational phases were calculated for each initial relative translational energy. The agreement for the reactive collisions is seen from Figs. 4 and 5 to be excellent.

Examination of Eqs. (1) and (2) shows that the use of natural collision coordinates introduces a singularity at $\kappa\rho=-1$, and indeed an ambiguity in the coordinate definition in this region.²⁴ If a trajectory approached

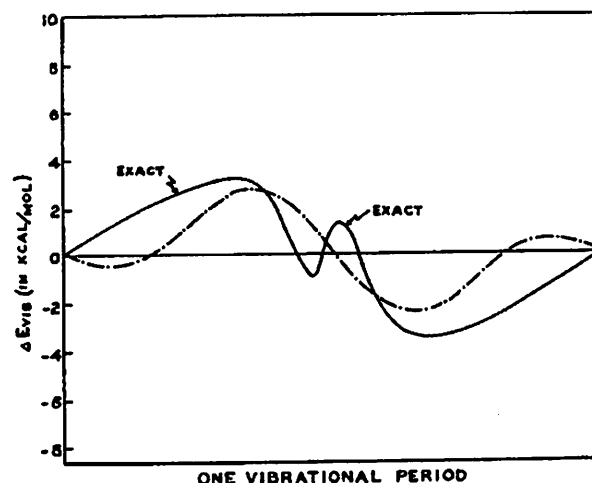


FIG. 6. Vibrational-translation energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{trans}^i=8.5$ kcal/mol.

this singularity or reached points corresponding to negative values of $1+\kappa\rho$, i.e., points on the other side of the singularity from curve C in Fig. 1, an error would result from the use of these natural collision coordinates. The excellent agreement of the two calculations for the reactive collisions shows that the singularity was not approached closely (at least for the cases treated in those figures). Figure 4 gives perhaps an indication of some difference for the nonreactive case, and indeed, as an examination of nonreactive trajectories will show later, that the system came closer to the singularity in this case.

C. Approximate Results Based on Eq. (14)

The adiabatic solution for $s(t)$ satisfies²

$$\dot{s} = [2\alpha(s)/\mu]^{1/2} (1 + \kappa\rho_0)^{-2}, \quad (17)$$

where $\alpha(s)$ is given by (6). Instead of $s(t)$ the co-

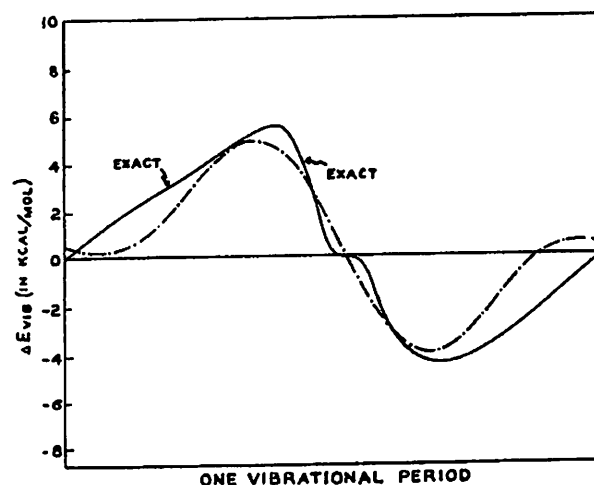


FIG. 7. Vibrational-translation energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{trans}^i=9.5$ kcal/mol.

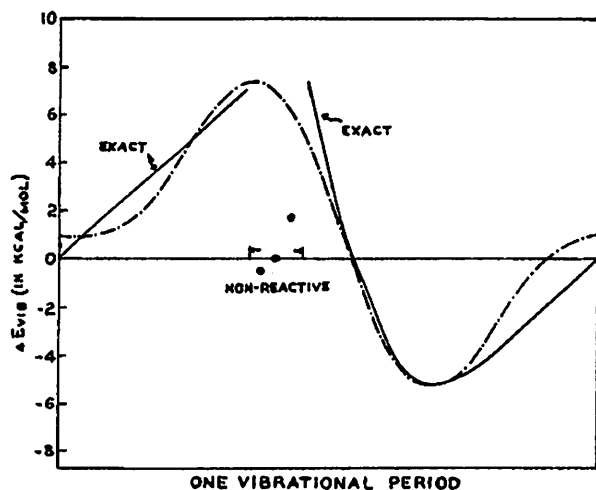


FIG. 8. Vibrational-translational energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{\text{trans}}^i = 10.55$ kcal/mol. The circled points refer to nonreactive collisions.

ordinate $v(t)$ defined earlier can be used, in which case (17) would become

$$\dot{v} = [2\alpha(s)/\mu]^{1/2} [1 + (du/dv)^2]^{-1/2} (1 + \kappa\rho_0)^{-2}. \quad (18)$$

While there was little advantage in using $v(t)$ instead of $s(t)$, provided the line $s=0$ was chosen to lie along the symmetry axis L of Fig. 1, i.e., at $v=0$, (18) was used instead of (17).

The phase δ_p in (14) is the vibrational phase when the system crosses the symmetry axis L in Fig. 1, i.e., the activated complex line. The trigonometric terms in (14) and (15) can be rewritten using the addition formulas, so that the phase δ_p appears outside each integral. Integrals denoted by S_1 , C_1 , S_2 , and C_2 may be defined:

$$S_1 = \int_{-\tau_0}^0 (\mu\omega\pi)^{1/2} \sin 2\pi w_p' dt,$$

$$C_1 = \int_{-\tau_0}^0 (\mu\omega\pi)^{1/2} \cos 2\pi w_p' dt,$$

$$S_2 = \int_{-\tau_0}^0 \frac{\dot{\omega}}{\omega} \sin 4\pi w_p' dt,$$

$$C_2 = \int_{-\tau_0}^0 \frac{\dot{\omega}}{\omega} \cos 4\pi w_p' dt, \quad (19)$$

where $-\tau_0$ is the initial time, i.e., the time when $v = -v_i$, and where the w_p' in (19) denotes

$$2\pi w_p' = \int_0^t \omega dt. \quad (19')$$

J_p at the (truncated) final time ($t = \tau_0$) is found from (14) to be

$$J_p(t = \tau_0) = J_p^0 - 4(J_p^0)^{1/2} S_1 \cos \delta_p + 4S_1^2 + 2J_p^0 C_2 \sin 2\delta_p. \quad (20)$$

J_p at the saddle point, $J_p(t=0)$, is similarly found to be $J_p(t=0) = J_p^0 - 2(J_p^0)^{1/2} (S_1 \cos \delta_p + C_1 \sin \delta_p) + S_1^2 + C_1^2 + (C_2 \sin 2\delta_p - S_2 \cos 2\delta_p) J_p^0$. (21)

The δ_p^0 used to represent the approximate results in Figs. 6–11 is related to δ_p by

$$\delta_p^0 = \delta_p - \int_{-\tau_0}^0 \omega dt. \quad (22)$$

The integrals in (19') and (22) were computed using a six-point Gauss-Legendre quadrature for an integration interval, $\Delta v = 0.005$ a.u. The integrals S_1 , C_1 , S_2 , and C_2 in (19) were next calculated by Simpson's rule. The value of J_p was then obtained from (20) and (21) at the product and saddle-point regions, respectively. The integrals are tabulated in Table III.

The change of vibrational energy ΔE_{vib} was calculated both from the exact Hamilton's equations of motion and from the above equations for a series of initial relative translational energies, $E_{\text{trans}}^i = 8.5$ –15.5 kcal/mol, and initial vibrational phases.²⁵ The results are given in Figs. 6–11.

Equations (6) and (17) yield the adiabatic threshold for reaction. However, for certain δ_p , the actual vibrational energy $E_{\text{vib}}(s, \delta_p)$ may exceed $E_{\text{vib}}(s)$ and so cause reflection at energies above the adiabatic threshold. Reflection occurs when the actual $\dot{s}$ vanishes, namely, when

$$E - E_{\text{vib}}(s, \delta_p) - V_1(s) - V_2(\rho_0, s) \leq 0 \quad (23)$$

at some s . Equation (23) can be solved for δ_p and for s , which will be near $s=0$ since $V_1(s)$ has a maximum there. If $s=0$ is used in (23),²⁶ and the latter is solved for δ_p , the nonreactive collisions so defined are indicated by vertical lines on the abscissas in Figs. 6–11. The range of nonreactive phase is seen to increase, and therefore, the reaction probability decreases, with increasing initial relative translational energy, at the higher energies, in the energy range considered.

Comparing the reactive portion in Figs. 6–11, we see that there is good agreement between the exact and

TABLE III. Values of integrals, S_1 , C_1 , S_2 , and C_2 .

E_{trans}^i (kcal/mol)	S_1	C_1	S_2	C_2
7.5	0.005330	-0.037280	0.080769	0.088301
8.5	-0.014951	-0.049132	0.093976	0.126385
9.5	-0.033779	-0.056788	0.103261	0.164801
10.55	-0.052518	-0.062774	0.112568	0.204917
11.5	-0.068083	-0.068069	0.112378	0.229329
12.5	-0.084476	-0.072271	0.115724	0.255077
13.5	-0.099598	-0.076646	0.117835	0.283105
14.5	-0.114736	-0.079938	0.121419	0.303617
15.5	-0.128878	-0.083120	0.121589	0.327182
16.5	-0.142668	-0.086128	0.125010	0.347025
17.5	-0.156488	-0.088406	0.120778	0.370785
19.5	-0.180849	-0.093940	0.128212	0.400339

approximate calculations of the phase-dependent ΔE_{vib} for most of the reactive regions. (In some instances the approximate curve sinusoidally winds around that of exact calculations in these figures, due to the character of the approximate Fourier series expansion employed.)

The phase-averaged vibrational energy change $\langle \Delta E_{\text{vib}} \rangle$ accompanying the reaction is plotted for a series of E_{trans}^i in Fig. 12. The vibrational phase-averaging procedure for the approximate calculation is merely the phase average over $\cos \delta_p$ and $\sin \delta_p$ in Eqs. (20) and (21). The average is over the range of reactive δ_p 's, with δ_p being random in this region. In Fig. 12 it is seen that the agreement between exact and approximate calculations of $\langle \Delta E_{\text{vib}} \rangle$ is good on an absolute basis.

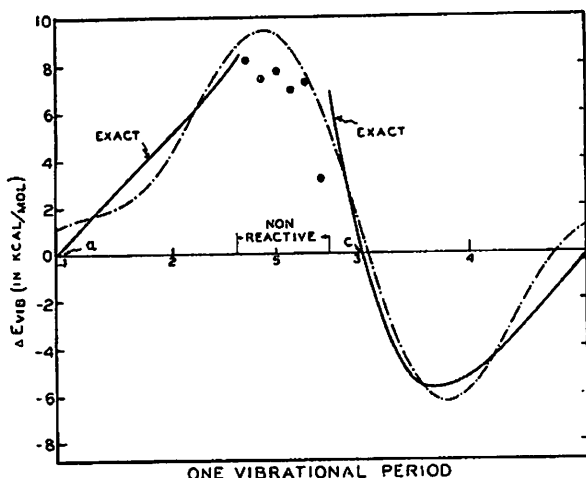


FIG. 9. Vibrational-translational energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{\text{trans}}^i = 11.5$ kcal/mol. The circled points refer to nonreactive collisions.

$\langle \Delta E_{\text{vib}} \rangle$ is seen in Fig. 12 to be small at initial relative translational energies appropriate to the usual reaction temperatures of 500–1200°K. (It can be shown that this range corresponds to E_{trans}^i in the neighborhood of 8.2–10.3 kcal/mol.) Implications for activated complex theory are noted later.

The maximum $\dot{\rho}_0$ contribution and the maximum $\dot{\omega}$ contribution to (14), obtained by maximizing each contribution with respect to δ_p , are plotted in Figs. 13 and 14. The $\dot{\rho}$ term and the $\dot{\omega}$ term are roughly equally effective in introducing nonadiabatic transitions at the lower energies. At the higher energies, the $\dot{\rho}$ term is larger than the $\dot{\omega}$ term (and the same remark applies to the phase-dependent $\dot{\rho}$ term alone). This behavior supports the treatment which led² to (14), since at low energies both terms need be treated only to first order while at higher energies the larger ($\dot{\rho}_0$) contribution should be treated to higher order.

The reaction probability for the exact classical calculation is

$$\text{Reaction Probability} = (\Delta \delta_p^0)_{\text{reactive}} / 2\pi, \quad (24)$$

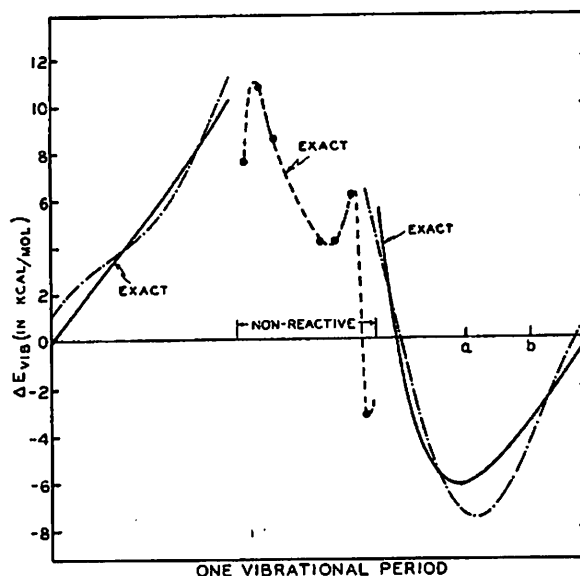


FIG. 10. Vibrational-translational energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{\text{trans}}^i = 13.5$ kcal/mol. The circled points and dashed line refer to non-reactive collisions.

where $(\Delta \delta_p^0)_{\text{reactive}}$ is the sum of intervals of initial vibrational phases for which the classical trajectories are reactive. The reaction probability is plotted versus initial translational energy in Fig. 15 and compared with one obtained using the approximate equations and the condition (23) at $s=0$.²⁷ The reaction probability is seen in Fig. 15 to rise rapidly to unity at energies above threshold and to provide a transmission at energies below the adiabatic threshold, i.e., to provide a "non-

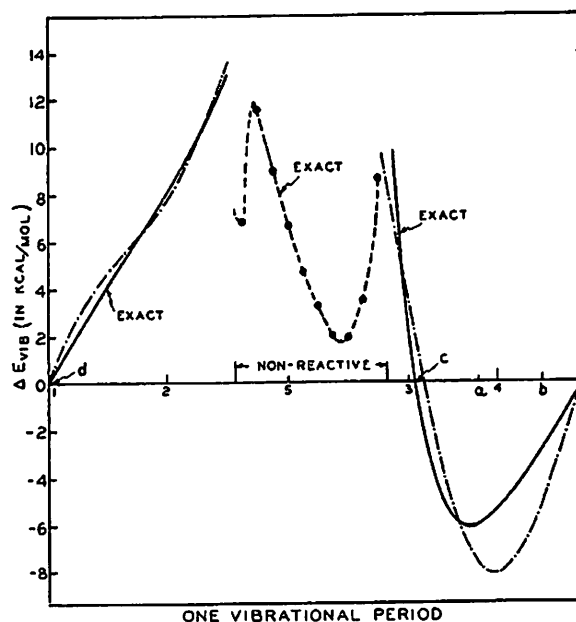


FIG. 11. Vibrational-translational energy transfer in a linear collision, as a function of initial vibrational phase: —, exact classical calculation; ---, approximate calculation, for $E_{\text{trans}}^i = 15.5$ kcal/mol. The circled points and dashed line refer to non-reactive collisions.

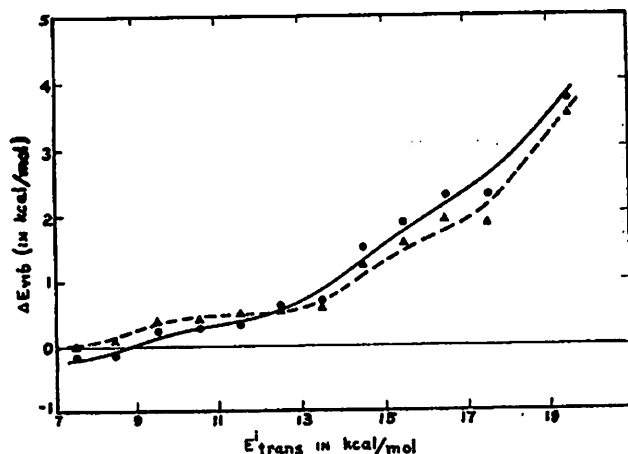


FIG. 12. Phase average of the vibrational-translational energy transfer for linear collision, as a function of initial relative translational energy: —●—, exact; —▲—, approximate.

adiabatic leak.” This same leak is responsible for the negative values of exact $\langle \Delta E_{vib} \rangle$ at low initial translational energies in Fig. 12.

The vibrational action computed from Eq. (14)²⁸ is given as a function of the coordinate s in Figs. 16 and 17 for two initial translational energies and various vibrational phases. Several not unexpected points emerge in these figures: (1) The rate of change of J_p along s is least at large $|s|$, where there is no interaction, and at small s , i.e., in the vicinity of the saddle point, where the velocity $\dot{s}$ along the reaction coordinate is least. (2) The region where $E - E_{vib}(s, \delta_p) - V_1(s) - V_2(\rho_0, s)$ becomes negative probably does not necessarily occur at $s=0$, although $V_1(s)$ is a maximum there, since $E_{vib}(s, \delta_p) [= J_p(s)\omega(s)/2\pi]$ sometimes has a local minimum at or near $s=0$. (3) Because of the increase in $J_p(s)\omega(s)$ after passage through $s=0$ in

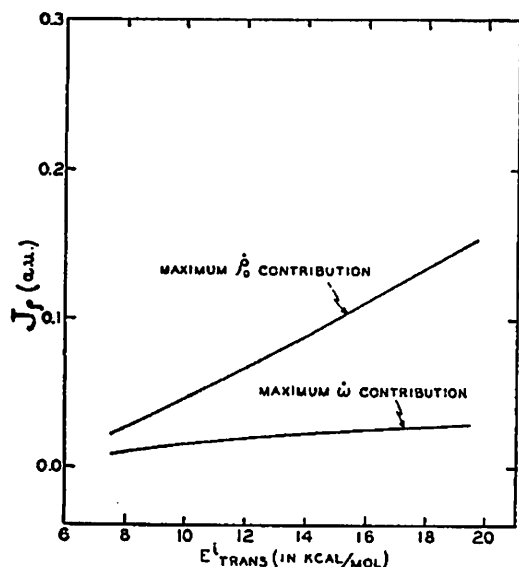


FIG. 13. Maximum contributions from $\dot{\rho}_0$ and ω terms to the action J_p at the saddle-point region (maximum over-all vibrational phases) as a function of initial relative translational energies.

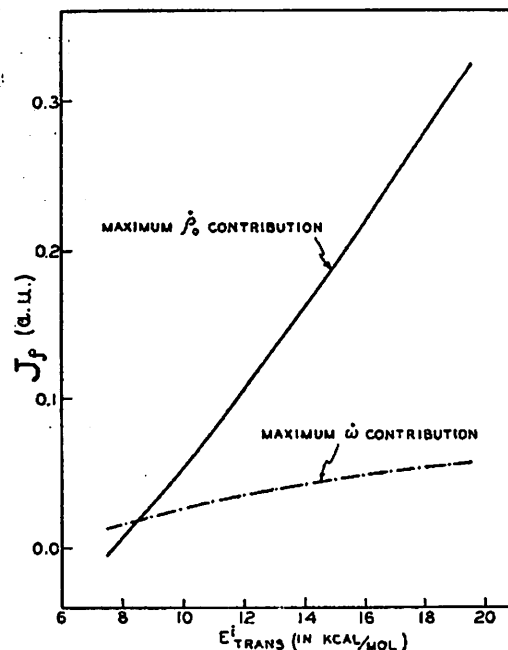


FIG. 14. Maximum contributions from $\dot{\rho}_0$ and ω terms to the final action J_p (maximum over-all vibrational phases) as a function of initial relative translational energies.

some instances, reflection and hence nonreactivity occurs after the system has passed through the saddle-point region. This reflection, involving as it does a re-crossing of line L , leads to a decrease in validity of activated complex theory at high energies.

To explore the last and other points a number of trajectories were plotted in Figs. 18 and 19. For these collisions, which have more than sufficient translational energy to overcome the adiabatic threshold, the non-reactivity is seen to occur as a result of reflection after the system has crossed the activated complex line L ($s=0$). Several other observations from these figures may also be made:

(1) There are two trajectories (No. 1 and one near No. 3, hereinafter referred to as No. 3) which are symmetric about the activated complex line L , and for these, of course, $\dot{\rho}=0$ and $\dot{\rho}-\dot{\rho}_0=0$ on line L . Thus,

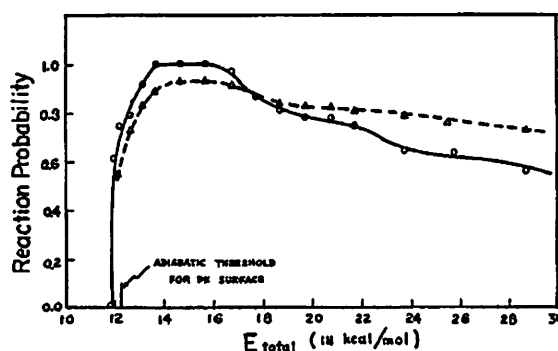


FIG. 15. Reaction probability for linear collision as a function of total energy: —, exact classical calculation; ---, approximate calculation based on Eqs. (23) and (14).

such trajectories cross line L at a turning point of vibrational ρ motion. These points lie to the left of the saddle point, as expected from the internal centrifugal force term α/η^2 . According to Figs. 18 and 19 for the exact trajectories, E_{vib} at line L and J_ρ at line L should be less for the symmetric trajectory 1 than for 3 since the vibrational potential energy is smaller in the former case and both have zero vibrational kinetic energy. This observation is also in agreement with the results for the approximate J_ρ 's based on Eq. (14) and given in Figs. 16 and 17, where trajectory 1 has a smaller J_ρ at $s=0$ than does trajectory 3.

(2) All the reactive trajectories were found to cross line L in an interval (a, b) , which is roughly bisected by the reaction coordinate. Thus, the ensemble of systems crossing L are more "centered" on the reaction

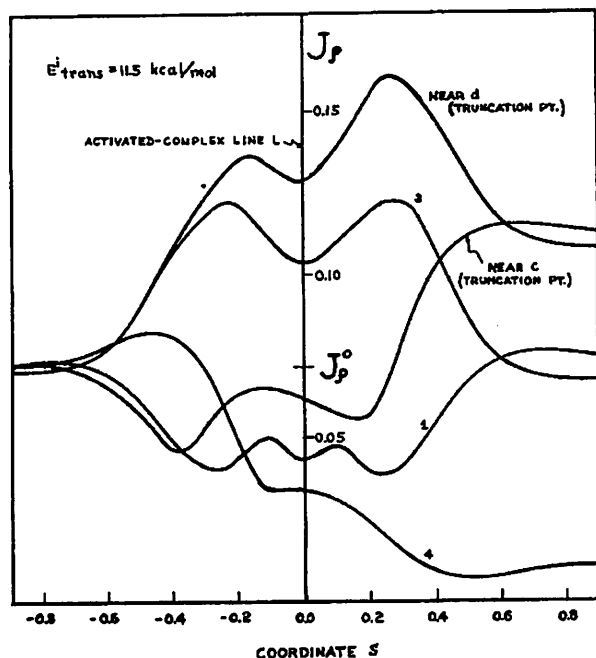


FIG. 16. Time evolution of the vibrational action $J_\rho(s)$ in a linear collision for $E_{trans} = 11.5$ kcal/mol. The corresponding initial vibrational phases are indicated in Fig. 9.

coordinate than on the curve of steepest descent (the "reaction path"), curve C in Fig. 1. Thereby, as predicted in Ref. 1, the use of a reaction coordinate which includes the dynamical curvilinear effect, as in the present case, provides a better zeroth-order separability between the vibrational and translational motions of the system than the use of the "reaction path" alone.

An effort was made to see if even better agreement of approximate and exact calculations in Figs. (6)–(11) could be made by integrating the approximate Eqs. (13) exactly.²⁹ Agreement was somewhat better in some regions, somewhat worse in others, and that more hybrid procedure was not explored further.

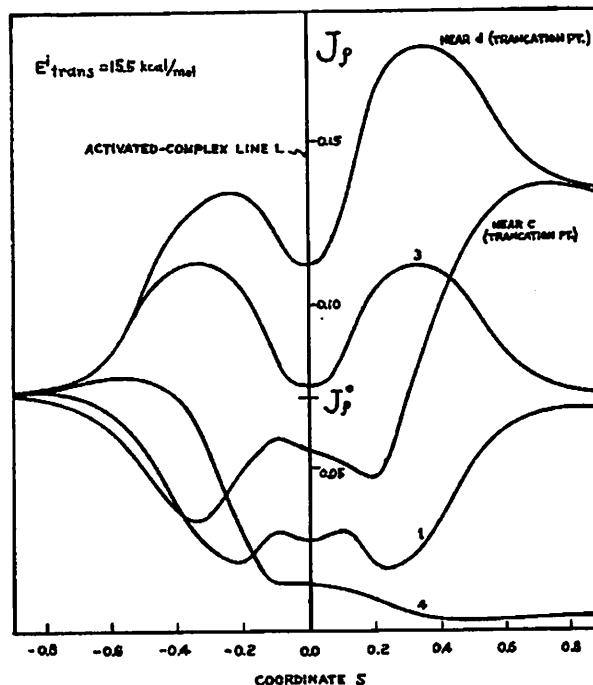


FIG. 17. Time evolution of the vibrational action $J_\rho(s)$ in a linear collision for $E_{trans} = 15.5$ kcal/mol. The corresponding initial vibrational phases are indicated in Fig. 11.

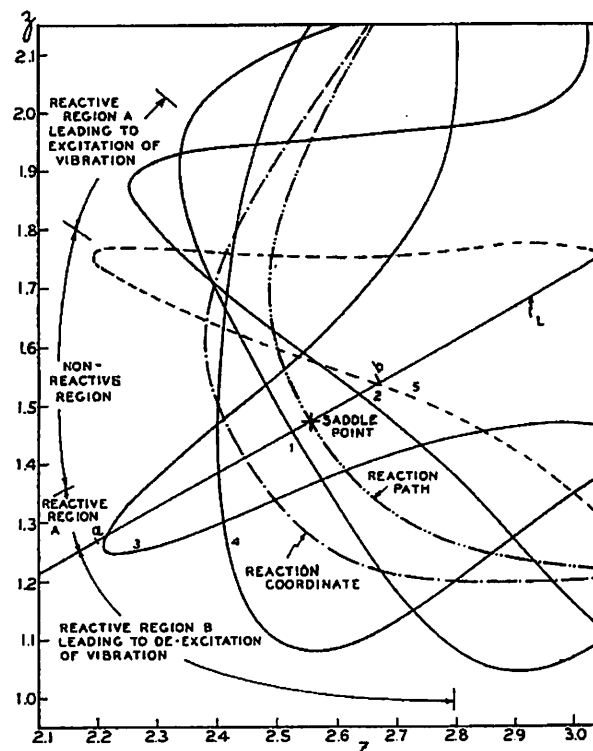


FIG. 18. Plot of classical trajectories on the potential energy surface near the saddle-point region, for the five vibrational phases indicated in Fig. 9: —, reactive trajectory; ---, non-reactive trajectory. ($E_{trans} = 11.5$ kcal/mol.) Curve C (---) and the reaction coordinate (---) are also plotted.

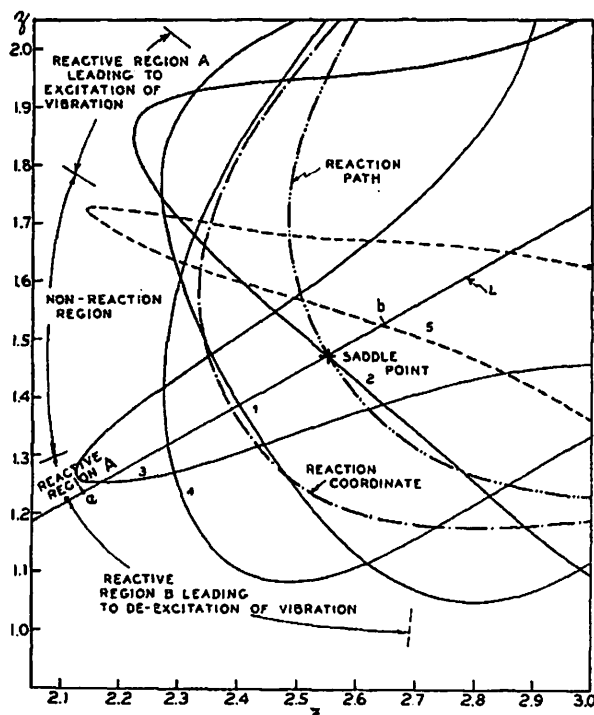


FIG. 19. Plot of classical trajectories on the potential energy surface near the saddle-point region, for the five vibrational phases indicated in Fig. 11: —, reactive trajectory; ---, non-reactive trajectory. ($E_{\text{trans}} = 15.5$ kcal/mol). Curve C (---) and the reaction coordinate (---) are also plotted.

CONCLUSIONS

Application of the Hamilton-Jacobi equation in the near-adiabatic approximation, and of local action-angle variables, was made to the linear reactive $\text{H}+\text{H}_2$ system. Good agreement was found between the exact and approximate calculation for the absolute value of the vibrationally phase-averaged vibrational energy change ($\langle \Delta E_{\text{vib}} \rangle$) accompanying reaction, in the region investigated (E_{trans} from 7 to 20 kcal/mol). (At the usual reaction temperatures of 500–1200°C, the E_{trans} contributing significantly to the reaction collisions on this surface would be in the neighborhood of 8–10.5 kcal/mol.) Good agreement for the phase-dependent ΔE_{vib} was also found for the data in Figs. 6–11. In effect, the curvilinear $[\rho_0]$ term in (14) was treated as a perturbation to second or higher order, while the $\dot{\omega}$ term was treated as a perturbation to first order, a method which later proved to be consistent with the findings in Figs. 13 and 14. Again, the phase-averaged vibrational energy change, $\langle \Delta E_{\text{vib}} \rangle$ in Fig. 12, was small for E_{trans} in the range appropriate to the usual reaction temperatures of 500–1200°K.

The centering of the reactive trajectories on the reaction coordinate defined by Eq. (7), as those trajectories crossed line L , lends further support to the concepts employed in (7)–(10). In the corresponding quantum mechanical treatment, a more truncated vibrational basis set can be used at each s if that set is centered upon a mean reaction coordinate curve [de-

fined by some $\rho_0(s)$] as compared with centering it on the curve C in Fig. 1.

On the basis of these results one can expect that the corresponding quantum treatment will attain similar good agreement, at least when tunneling effects in the s motion can, as in the present classical case, be neglected. Tunneling can, incidentally, lead to one difficulty in the case of the present coordinates: It causes a "negative" internal centrifugal force^{1a} since $\alpha(s)$ in Eq. (6) is negative, and can thus cause the system to come near the singularity at $1+\kappa\rho=0$ if curve C is selected as in Fig. 1. In such cases it may be necessary to use as curve C a curve appreciably displaced from the steepest descent curve in Fig. 1, or to solve the problem in the presence of the singularity.

The findings of a small $\langle \Delta E_{\text{vib}} \rangle$ at low energies in Fig. 12 indicate that the reaction is statistically adiabatic²⁰ at the usual reaction energies, even though, as Figs. 6–11 show, it is not exactly adiabatic.²¹ Thus, upon energy averaging at a particular temperature, activated complex theory should be valid for describing the reaction rate in the classical system described in this paper. Deviations could nevertheless occur at low translational energies and hence at low temperatures because of the nonadiabatic leak, at least in the present case. Deviations would also occur at very high energies, and, thereby, at extremely high temperatures, due to the recrossing of the activated complex line L discussed in the preceding section.

This concept of adiabaticity has provided a zeroth-order basis for separating the motions perpendicular to the reaction coordinate from that parallel to that coordinate. This separation is the more valid the more slowly the molecular properties (such as the rate of change of vibrational frequency and the rate of change of tangent vector of the reaction curve) vary along that curve. Natural collision coordinates are particularly suited to such system, and hence to understanding activated complex theory. They should become less useful when the molecular properties vary rapidly along the curve C , e.g., when the curvature $\kappa(s)$ becomes very large. Such reactions are strongly non-adiabatic and are undoubtedly better treated by an impulse type of approximation.

* Acknowledgement is made to the donors of the Petroleum Research Fund, administered by the American Chemical Society, for partial support of this research. This research was also supported by a grant from the National Science Foundation at the University of Illinois.

¹ (a) R. A. Marcus, *J. Chem. Phys.* **45**, 4493 (1966), Paper I; (b) **45**, 4500 (1966), Paper II; (c) **49**, 2610 (1968), Paper III; (d) *Discussions Faraday Soc.* **44**, 7 (1967), and discussion remarks.

² R. A. Marcus, *J. Chem. Phys.* **49**, 2617 (1968), Paper IV. In Eq. (93) the $(2J_z)^{1/2}$ should read $2(J_z)^{1/2}$.

³ (a) D. Russell and J. C. Light, *J. Chem. Phys.* **51**, 1720 (1969); (b) C. C. Rankin and J. C. Light, *ibid.* **51**, 1701 (1969).

⁴ R. E. Wyatt, *J. Chem. Phys.* **51**, 3489 (1969).

⁵ J. N. L. Connor and M. S. Child, *Mol. Phys.* **18**, 653 (1970).

⁶ See, e.g., M. Born, *The Mechanics of the Atom* (Ungar, New York, 1960).

⁷ (a) A. O. Cohen and R. A. Marcus, *J. Chem. Phys.* **49**, 4509 (1968); (b) **52**, 3140 (1970).

⁸ M. Attermeyer and R. A. Marcus, *J. Chem. Phys.* **52**, 393 (1970).

⁹ Action-angle variables have also been applied to the calculation of transition probabilities by W. H. Miller (private communication) and by the author.

¹⁰ F. T. Wall, L. A. Hiller, Jr., and J. Mazur, *J. Chem. Phys.* **29**, 255 (1958); **35**, 1284 (1961); F. T. Wall and R. N. Porter, *ibid.* **39**, 3112 (1963).

¹¹ M. Karplus, R. N. Porter, and R. D. Sharma, *J. Chem. Phys.* **43**, 3259 (1965); **40**, 2033 (1964).

¹² (a) R. A. Marcus, *J. Chem. Phys.* **46**, 959 (1967); (b) Ref. 1(b).

¹³ E. M. Mortensen, *J. Chem. Phys.* **49**, 3526 (1968).

¹⁴ E. M. Mortensen and K. S. Pitzer, *Chem. Soc. (London), Spec. Publ.* **16**, 57 (1962); E. M. Mortensen, *J. Chem. Phys.* **48**, 4029 (1968).

¹⁵ D. Micha, *Arkiv Fysik* **30**, 425, 437 (1965); *J. Chem. Phys.* **50**, 722 (1969).

¹⁶ (a) M. Karplus and K. T. Tang, *Discussions Faraday Soc.* **44**, 56 (1967); (b) K. T. Tang and M. Karplus, *J. Chem. Phys.* **49**, 1676 (1968).

¹⁷ H. C. Corben and P. Stehle, *Classical Mechanics* (Wiley, New York, 1960).

¹⁸ α is the α_s of Ref. 2. We have, in effect, utilized Eqs. (11) (16), (81) and (82) of Ref. 2 to obtain Eq. (6), without necessarily employing the local-harmonic approximation used in (81). The ϵ in Eq. (81) is the present $E_{\text{vib}}(s)$.

¹⁹ We have chosen an arbitrary constant, t_0 , in (93) and (94) to be zero.

²⁰ See, for example, references cited in I. Shavitt, R. M. Stevens, F. L. Minn, and M. Karplus, *J. Chem. Phys.* **48**, 2700 (1968).

²¹ R. N. Porter and M. Karplus, *J. Chem. Phys.* **40**, 1105 (1964).

²² The equations of motion in Cartesian coordinates^{9,10} are $\dot{z} = p_z/\mu$, $\dot{p}_z = -\partial V/\partial z$, $\dot{Z} = p_Z/\mu$, $\dot{p}_Z = -\partial V/\partial Z$, where V is the potential energy. At end of the calculation z and Z are converted to Cartesian coordinates appropriate to products.

²³ The step size of the integration, Δt , was 0.02×10^{-14} sec, which was approximately 1/25 of the vibrational period of H_2 molecule. Computing time on the IBM 7094 for one initial relative translational energy with 36 vibrational phases was approximately 10 min with a step size of 0.02×10^{-14} sec; the exact value depended upon the initial relative velocity.

²⁴ Ref. 1(a), Footnote 16.

²⁵ Computing for a single translational energy and an arbitrary number of vibrational phases was 2 min.

²⁶ The rigorous solution for s in (23) is straightforward, but the needed integrals in (14) were not printed out.

²⁷ The more rigorous solution of (23)²⁸ would have lowered the transmissions at high E_{trans} , and so perhaps have improved the agreement in Fig. 15.

²⁸ Actually the results in these figures are obtained from the integration of the equations for the ϕ_1 and ϕ_2 functions in Ref. 29.

²⁹ For this purpose real combinations of the a and a^* were introduced, defined by $\phi_1 = J_p^{1/2} \cos 2\pi w_p$, $\phi_2 = J_p^{1/2} \sin 2\pi w_p$, and the resulting equations for ϕ_1 and ϕ_2 were integrated numerically.

³⁰ R. A. Marcus, *J. Chem. Phys.* **49**, 2617 (1968); Ref. 12(b).

³¹ For the relation of adiabatic theory to activated complex theory see, for example, M. A. Eliason and J. O. Hirschfelder, *J. Chem. Phys.* **30**, 1426 (1959), and other references in R. A. Marcus, *ibid.* **45**, 4493 (1966), and Appendix III of Ref. 11(a). Incidentally, an adiabaticity or statistical adiabaticity from reactants' region to activated complex region suffices to yield activated complex theory, provided no subsequent reflection of the system occurs during its motion from activated complex region to products region.